

Combined high field and benchtop NMR studies of catalyst separation techniques on proton exchange membranes

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Motivation

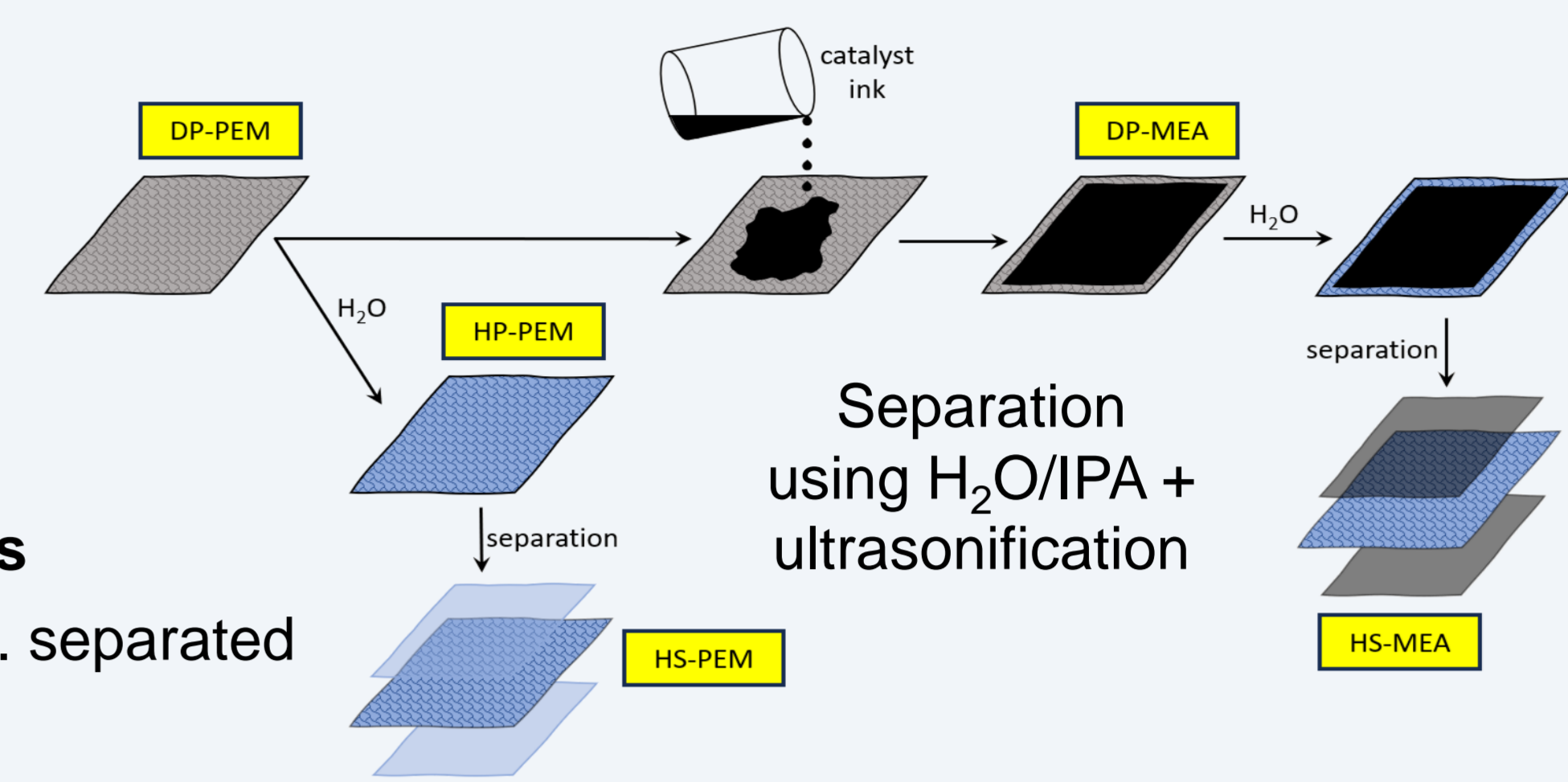
- Proton exchange membrane water electrolyzers (PEMWE) are a mature, but not yet widely deployed technology for green H₂ production
- Due to high CRM cost and potential PFAS restrictions, recycling (separation) of both catalysts and PEM in membrane-electrode-assemblies (MEAs) is incentivized
- Development of fast screening technique for industry

Experimental

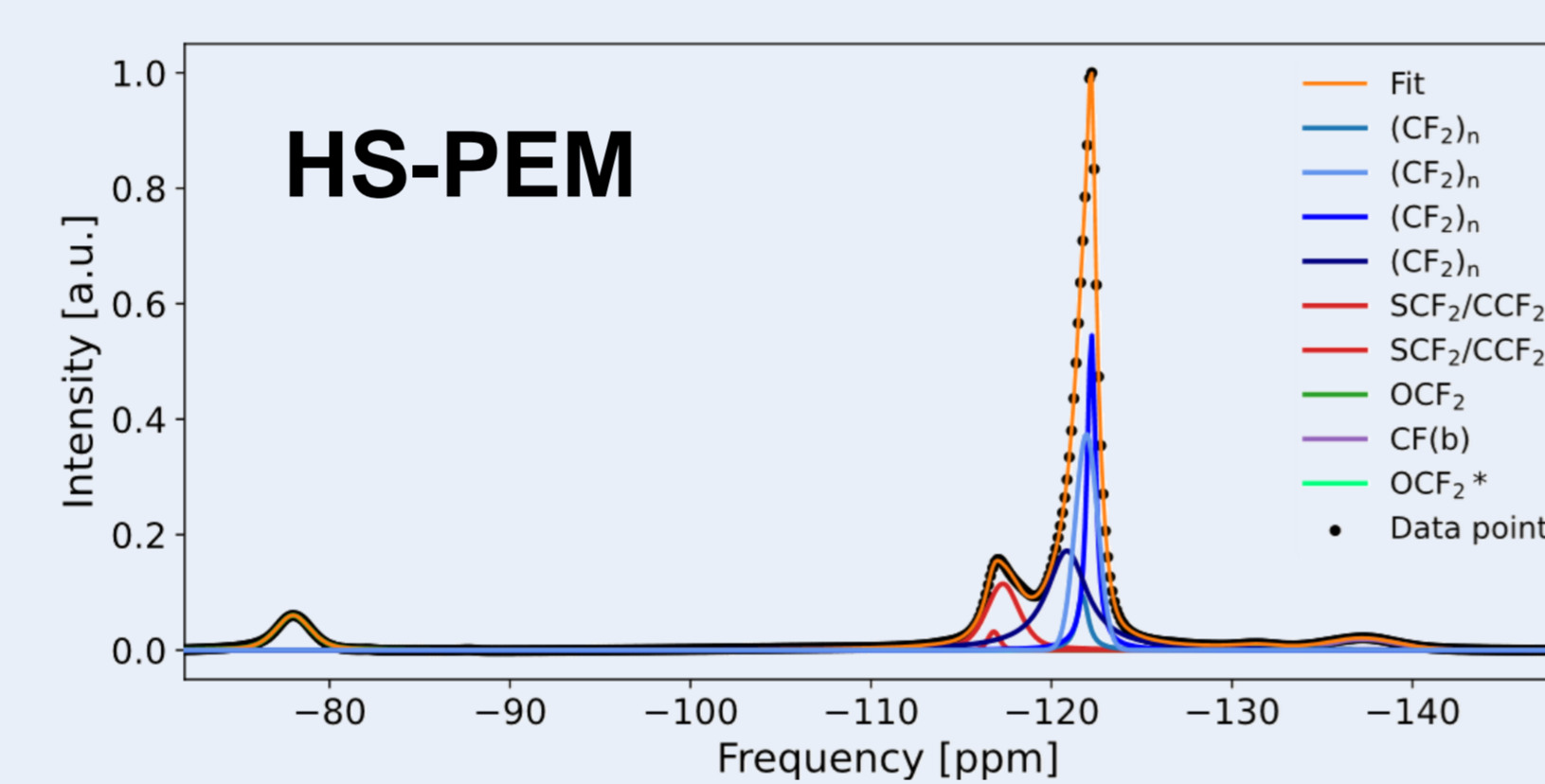
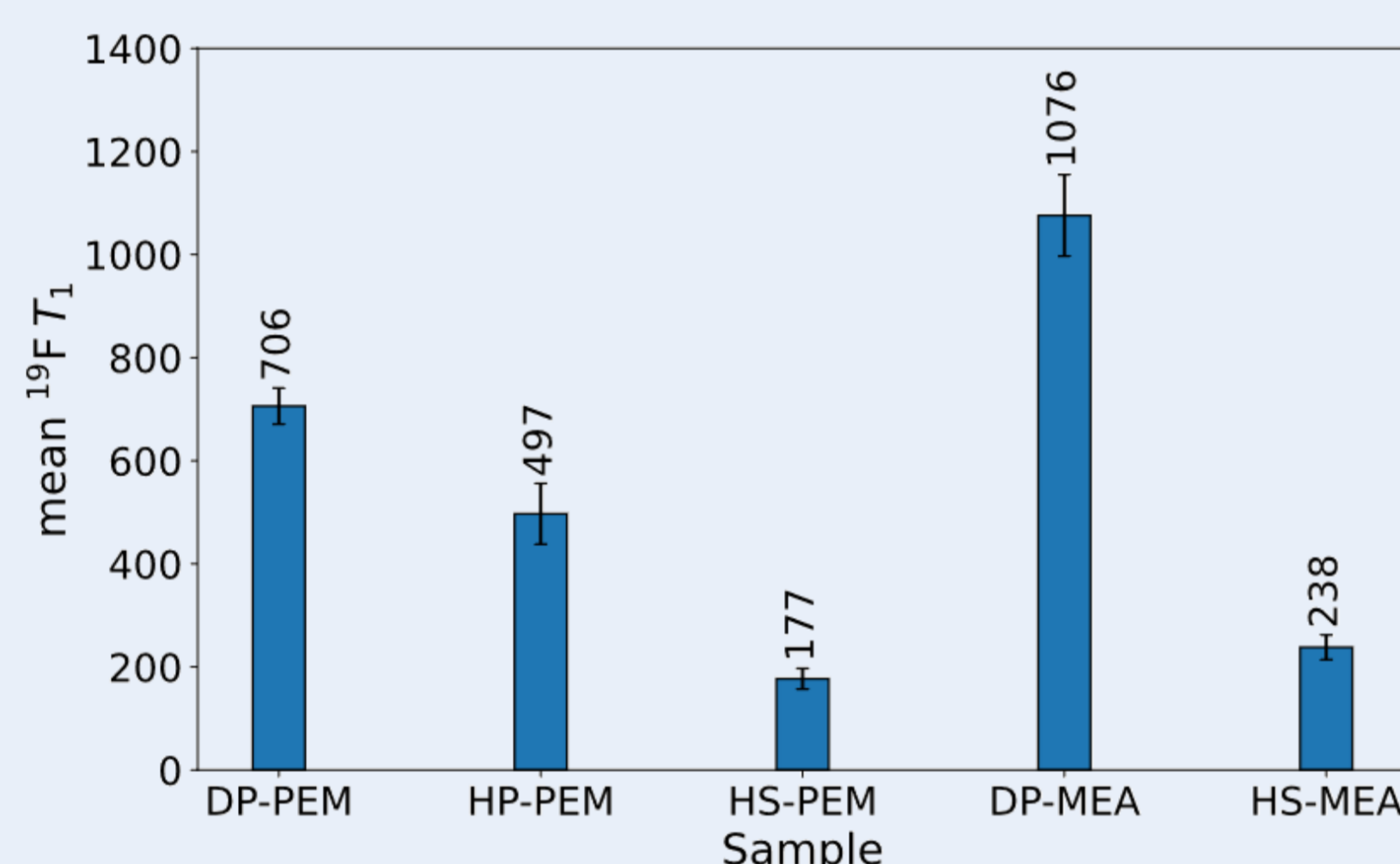
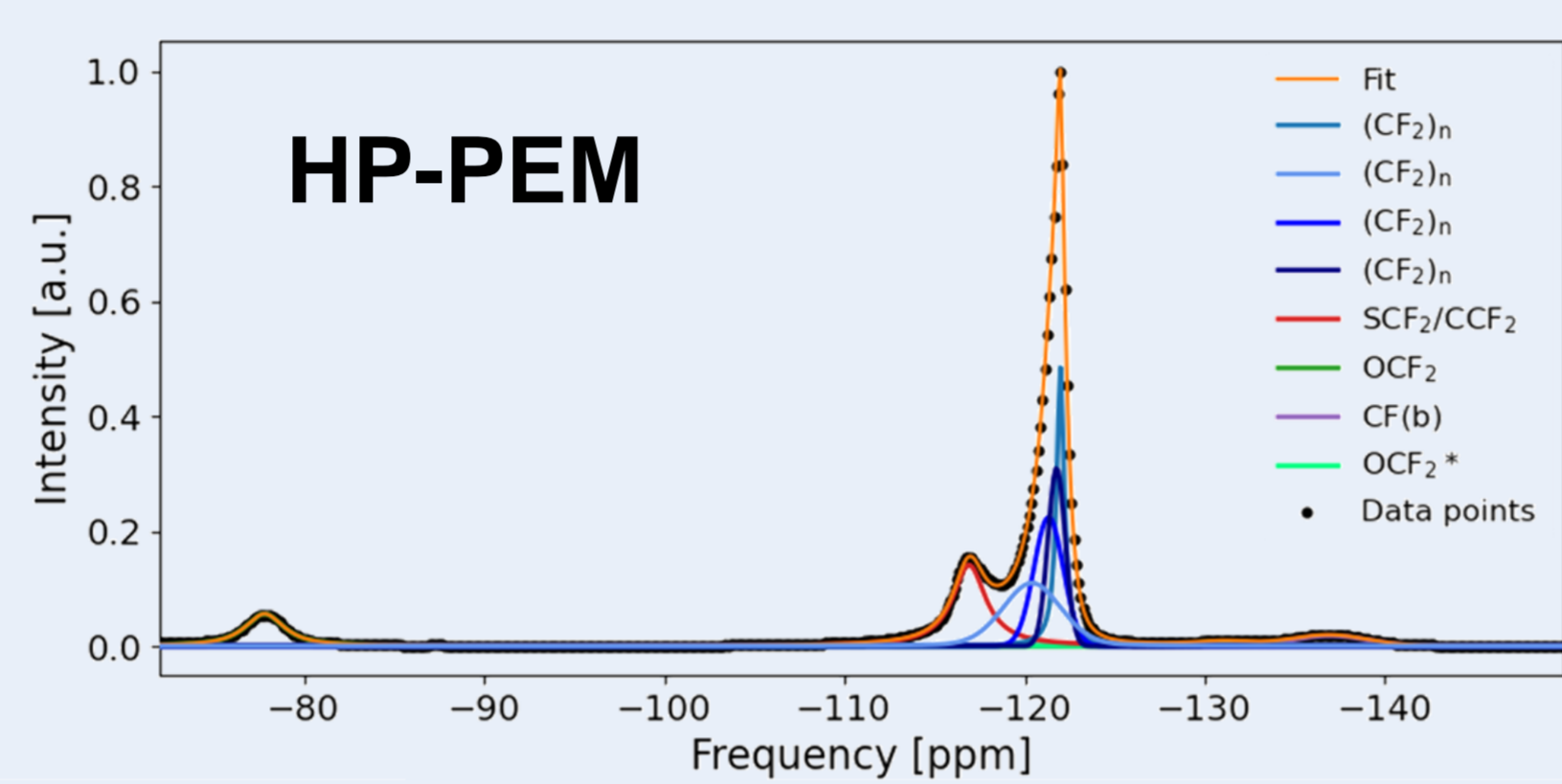
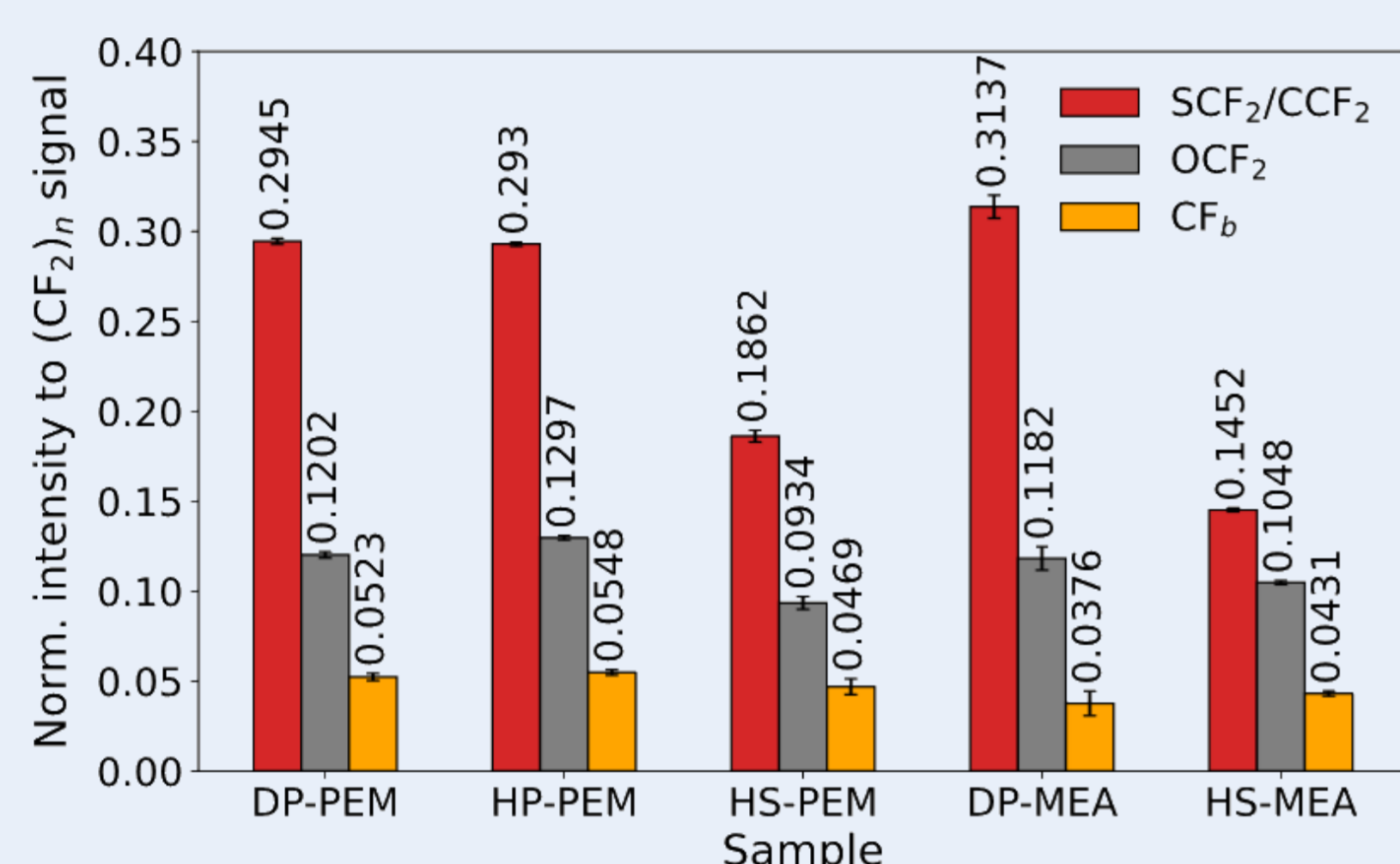
Dry
Hydrated
Pristine
Separated

Conditioning of PEMs

- Dry vs. hydrated vs. separated
- PEM vs. MEA



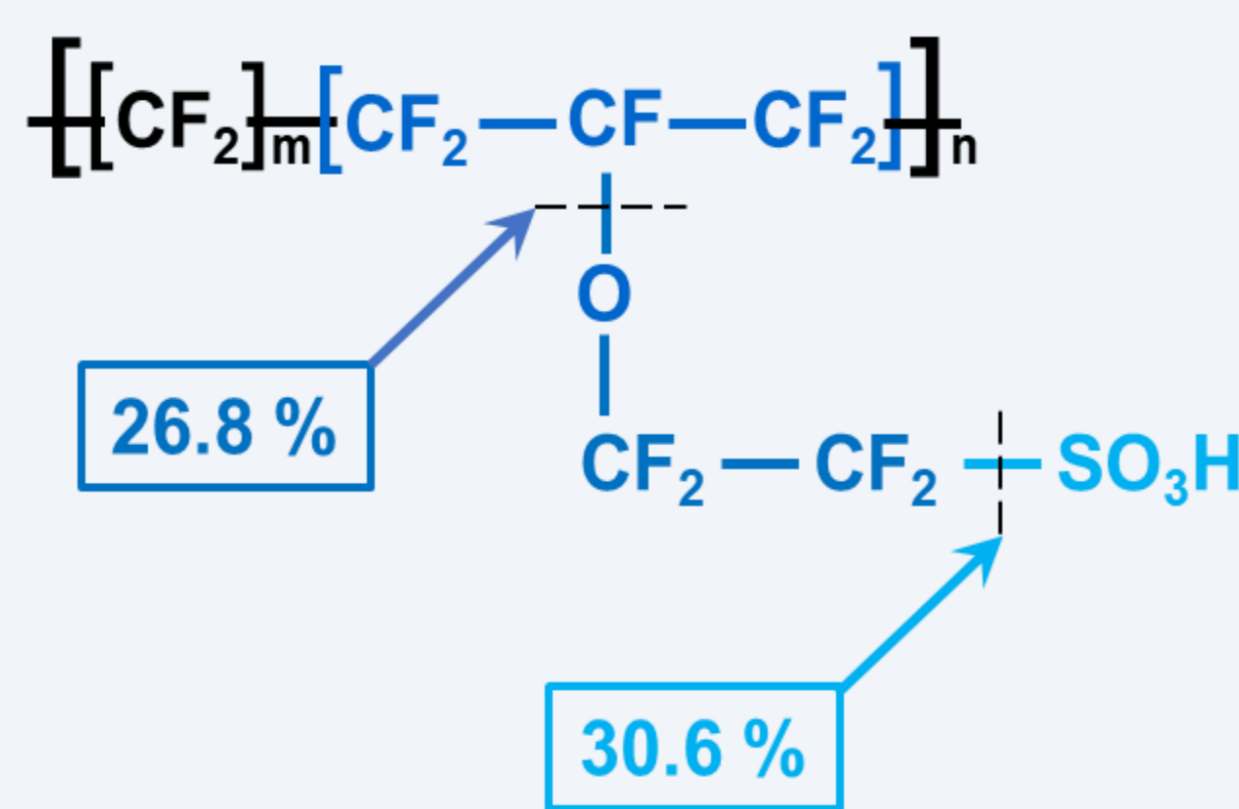
High-Field MAS NMR (400 MHz)



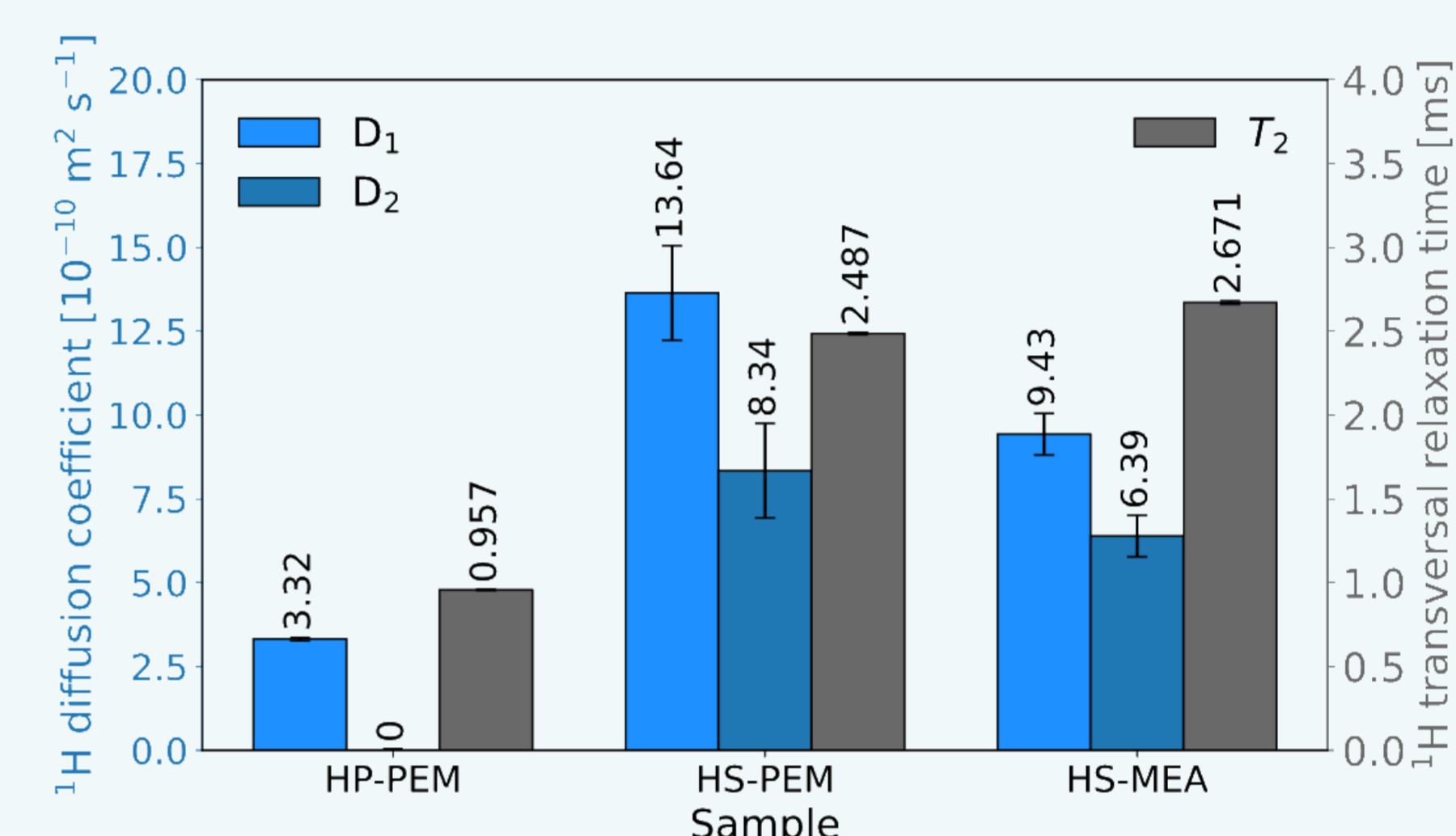
- Separation technique impacts chemical & spatial structure of ionomer
- ¹⁹F MAS spectra**
 - Loss of polar side chain after separation, particularly SO₃H group
 - Significant change in SCF₂ peak shape after separation
- ¹⁹F T₁ relaxation**
 - Strong spin diffusion for all samples
 - Significant decrease of T₁ after PEM hydration and separation

Discussion

- Separation technique induces chemical bond breakage in the polar side chain
- Side chain loss occurs via chain unzipping starting from SO₃H group,
 - C-S bond breakage as rate determining step
- Change in spatial structure due to side chain loss affects water uptake and transport behavior



Benchtop NMR (80 MHz)



- HP-PEM**
 - Single observable diffusion mode
 - Low-medium T₂
- HS-PEM/HS-MEA**
 - Significantly increased diffusivity
 - Two observable diffusion modes
 - Significantly increased T₂
- Correlation of proton relaxation time (T₂) and proton diffusivity (D₁, D₂)

Conclusions

- Separation using H₂O/IPA and ultrasonification causes non-negligible degradation of PEM properties via side-chain loss
- ¹⁹F high-field MAS NMR**
 - Precise study of PEM chemistry and structure
- ¹H benchtop NMR**
 - Correlative findings
 - Low cost and complexity
 - Predestined for industrial screening and quality control

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